CHEMICAL KINETIC SIMULATIONS

 $\mathbf{B}\mathbf{y}$

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WHAT IS CHEMICAL KINETICS?

- Chemical Kinetics is the study of the speed with which a chemical reaction occurs and the factors that affect this speed.
- Chemical Kinetics is useful to determine *how* a reaction occurs
- Study of the rate of reactions

RATE LAWS

- A Rate Law is an equation that tells us how fast the reaction proceeds and how the reaction rate depends on the concentrations of the chemical species involved.
- Common type of Rate Laws
 - Zero Order $r = -d[A]/dt = k[A]^0 = k$
 - First Order $r = k[A]^1$
 - Second Order $r = k[A]^2$

WHAT IS CONSECUTIVE FIRST ORDER REACTION

- Consecutive First Order Reaction is a chemical process in which the product of one reaction is the initial material for other reactions.
 - $A \rightarrow B \rightarrow C$
 - A turns into B, with rate constant k_1 , which turns into C, with rate constant k_2

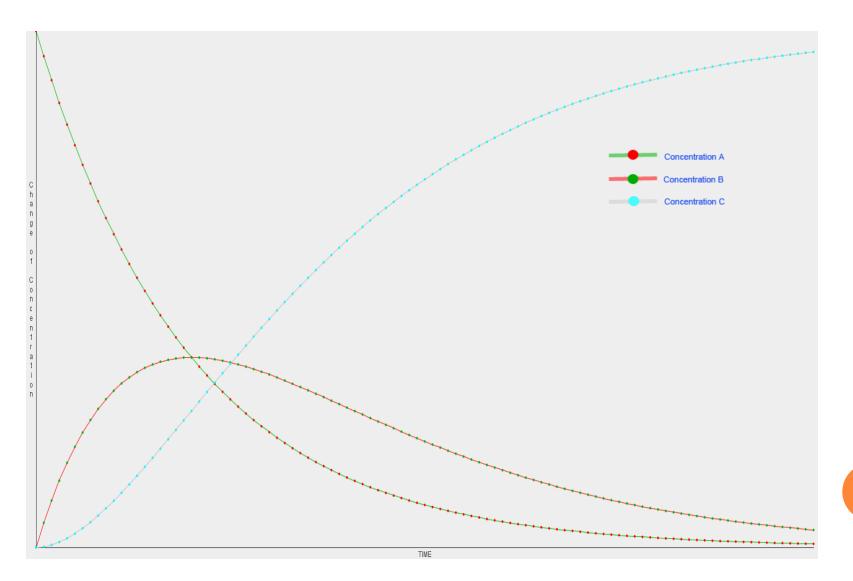
WHAT WE DID

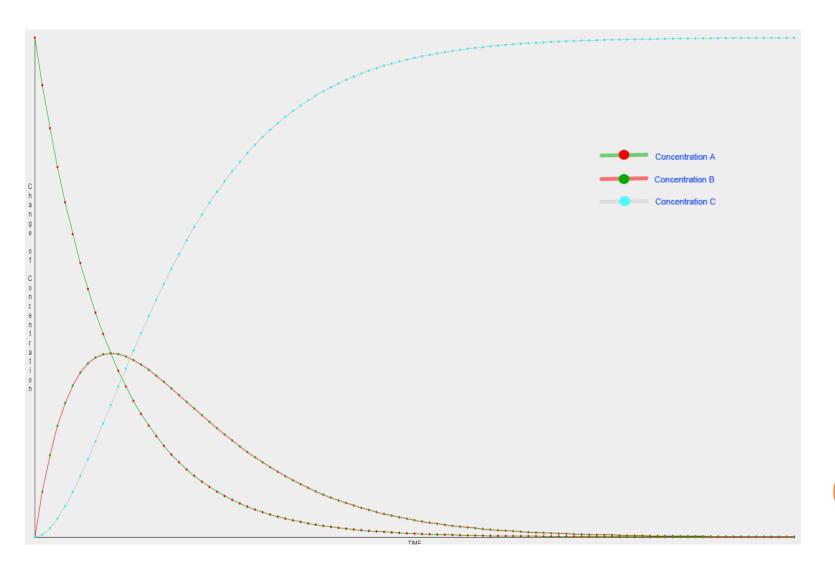
- Integrated the Rate Laws
 - In order to understand how the concentrations of the chemicals in a chemical reaction changes with time it is necessary to integrate the Rate Law.
 - $[A]_t = [A]_0 * e^{-k*t}$
 - $[B]t^{=}[A]_{0}$ *k*t*e-k*t+ $[B]_{0}$ *e-k*t
 - $[C]_t = [A]_0 * (1 e^{-kt} k * t * e^{-kt})^+ [B]_0 * (1 e^{-k * t})$
- Used the Integrated forms of Rate Laws for consecutive first order reaction A → B → C and plotted them using a Java Plotter Program.

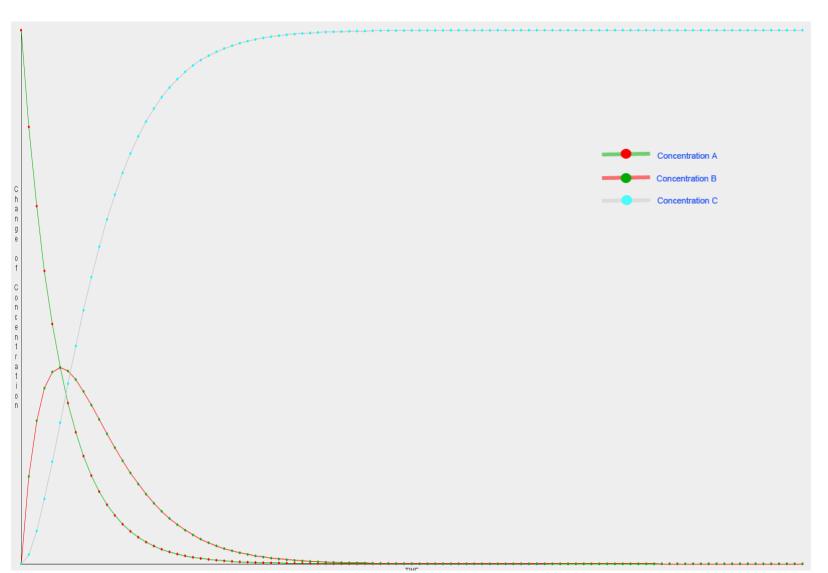
SIMULATIONS

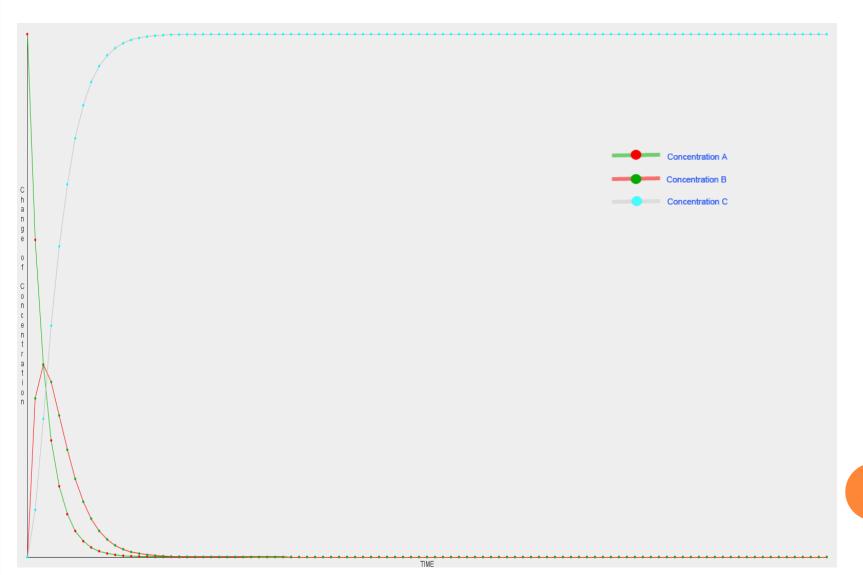
- Coded the Integrated Forms of the Rate Laws in Java.
- Ran Tests using the Integrated Forms of the Rate Laws with Different Constants of k.
 - Ranging from k = .5 to k = 50
- Gathered data from time interval t = .1 to t = 10
- Plotted 100 points for each chemical A, B, and C

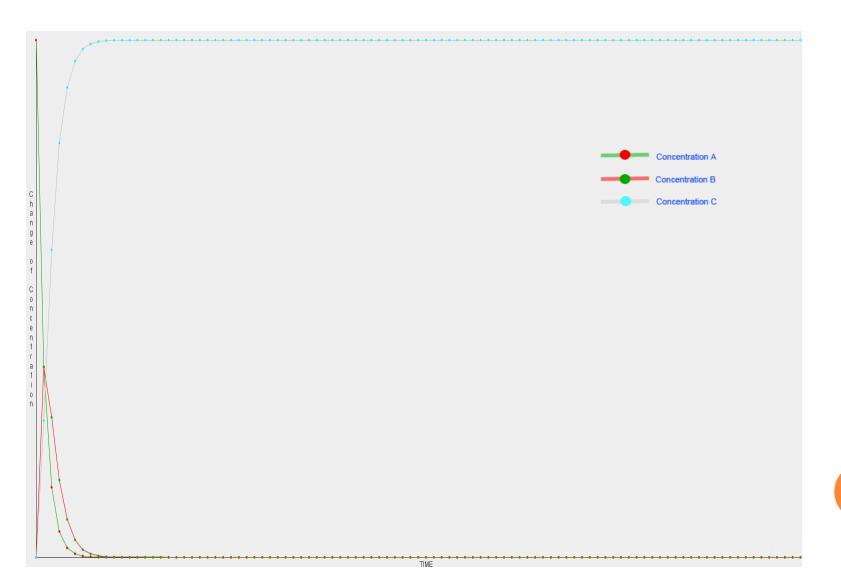
SIMULATIONS FOR A->B->C

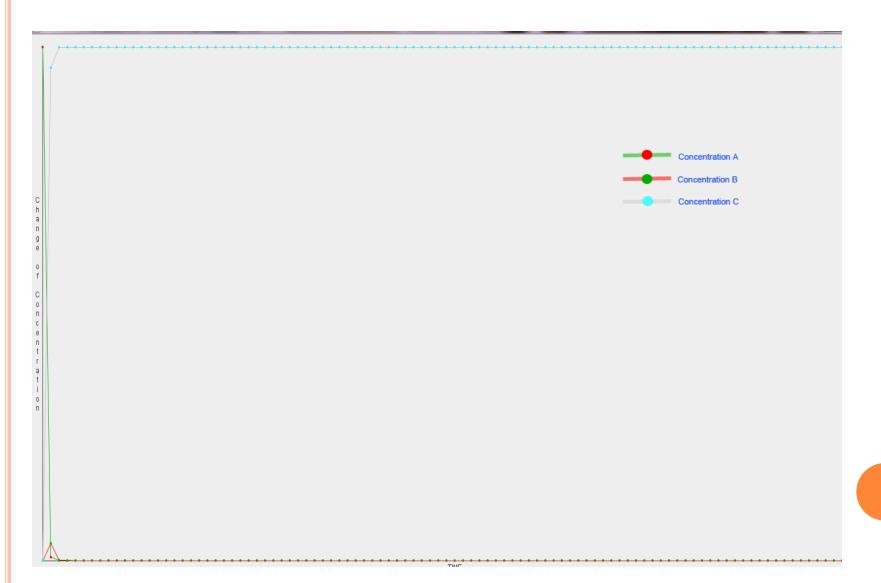




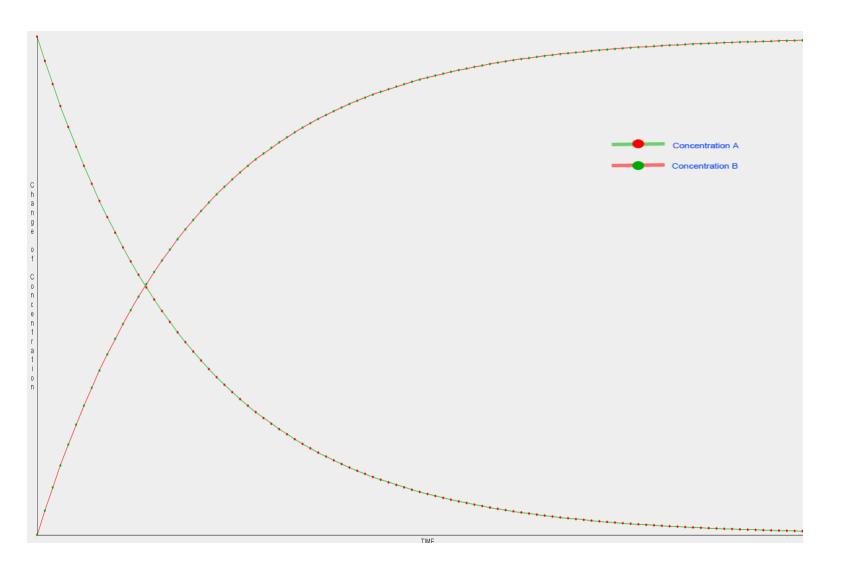


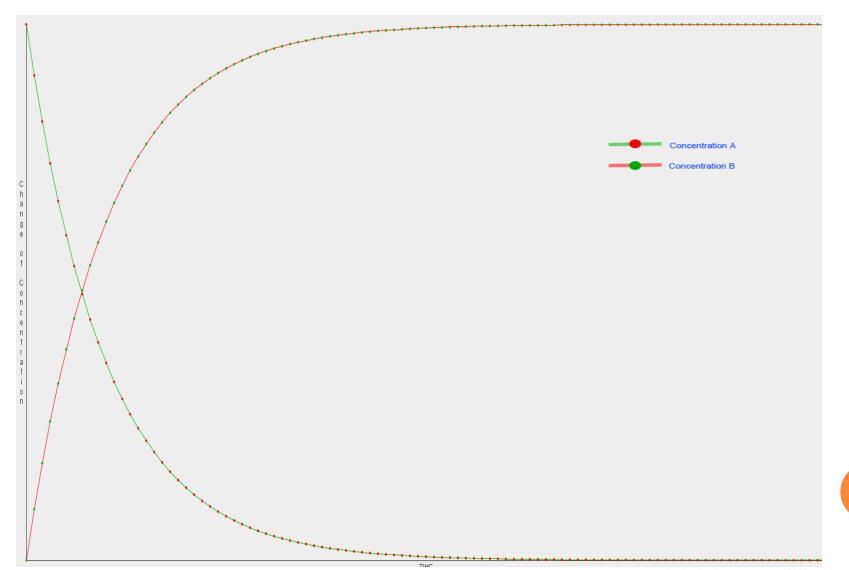


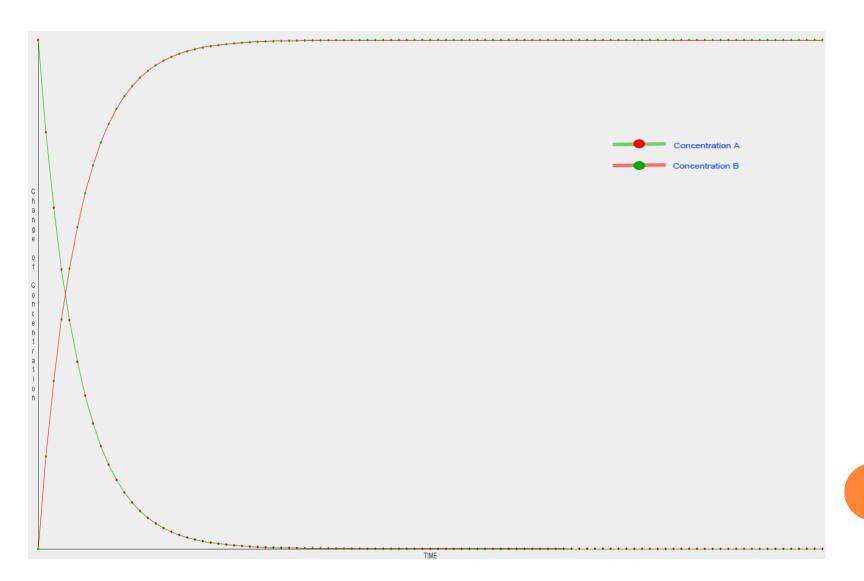


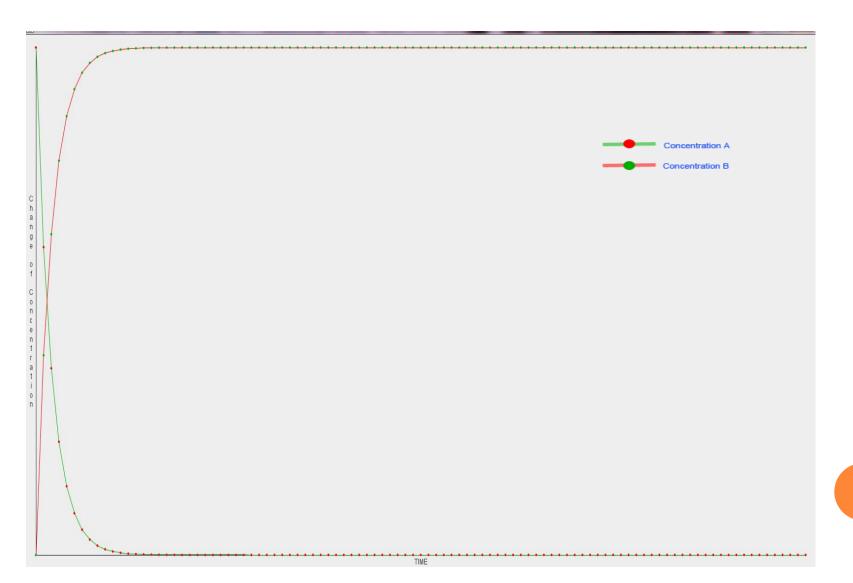


SIMULATIONS FOR A-B

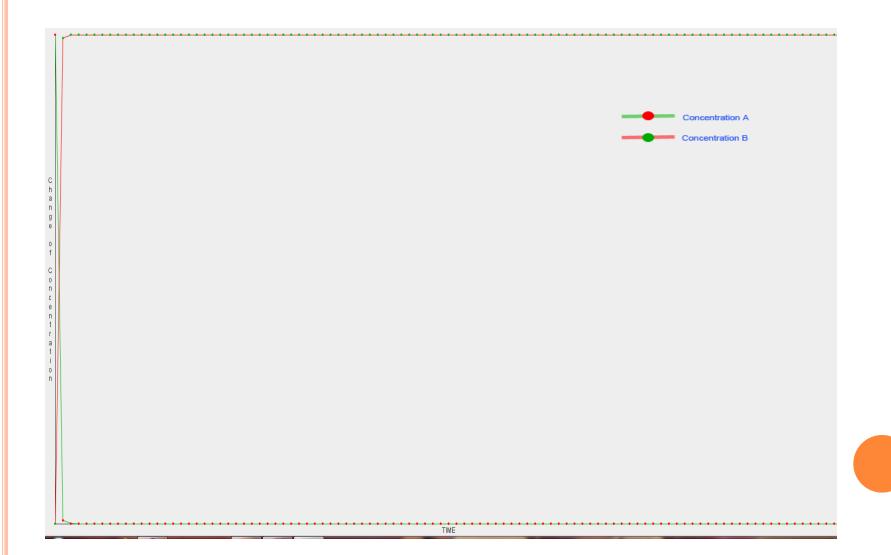












FUTURE WORK

- Run Simulations for $A \rightarrow B \rightarrow C \rightarrow D$
- Run Simulations for different values of k, l, and m

QUESTIONS

